**Directions:** Complete the following after watching the Part 5 Video.

1. How many protons, neutrons and electrons would be found in atoms as described by each isotopic notation given below?
2. Complete the chart.

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| **Isotopic Notation** | **Atomic Number** | **Mass Number** | **Number of Protons** | **Number of Electrons** | **Number of Neutrons** | **Neutral: What is the net charge?** |
|  |  |  |  |  |  |  |
| **Ca** |  | 41 |  | 18 |  |  |
|  |  |  | 50 | 46 | 50 |  |
|  | 11 | 25 |  | 10 |  |  |
|  |  | 130 | 52 |  |  | -2 |
| **F** |  |  |  |  | 10 | Neutral, 0 |

1. Write the complete isotopic notation for atoms matching each description. Assume that atom is neutral in each case.

a)     Contains 15 neutrons and 13 protons. d) Iron contains 84 subatomic particles.

b)    Atomic number is 41. e) Rubidium that contains 124 subatomic particles.

c)     Contains 33 electrons and 42 neutrons.